

Examination of the Biodiesel Production Performances of Natural and Modified Bentonite Heterogeneous Basic Catalysts

Sema ASLAN^{1*}

ABSTRACT: Intension of this study is the production of an alternative energy source from rapeseed oil with renewable heterogeneous basic catalysts. Production of biodiesel from biomass has been studied over decades and the positive effects of the heterogeneous basic catalysts on this process are still worth to examine. In the presented study firstly, natural bentonite (BT) was activated by calcination to obtain calcinated bentonite (C-BT). Then Li impregnation experiments were completed, and Li impregnated bentonite (Li-BT) catalyst was prepared. Obtained catalysts were characterized and analysed with FTIR, TGA and SEM/EDS measurements. Measurements showed that Li-BT is a promising catalyst and due to Li impregnation, exchange ability of the surface of the BT increased. Subsequently, they were utilized to the biodiesel production from rapeseed oil process individually and total fatty acid methyl ester yields of the obtained biodiesel were examined by GC-MS measurements. As a result, the best conversion yield was obtained from Li-BT used experiments as 98.80 % besides C-BT revealed with no yield. These studies can be enlarged by the consecutive catalytic experiments in detail. In conclusion, there is no need to use expensive catalysts and petroleum sourced hazardous chemicals to produce renewable energy catalysts.

Keywords: Heterogeneous basic catalyst, Biodiesel, Rapeseed oil, Bentonite, Renewable energy

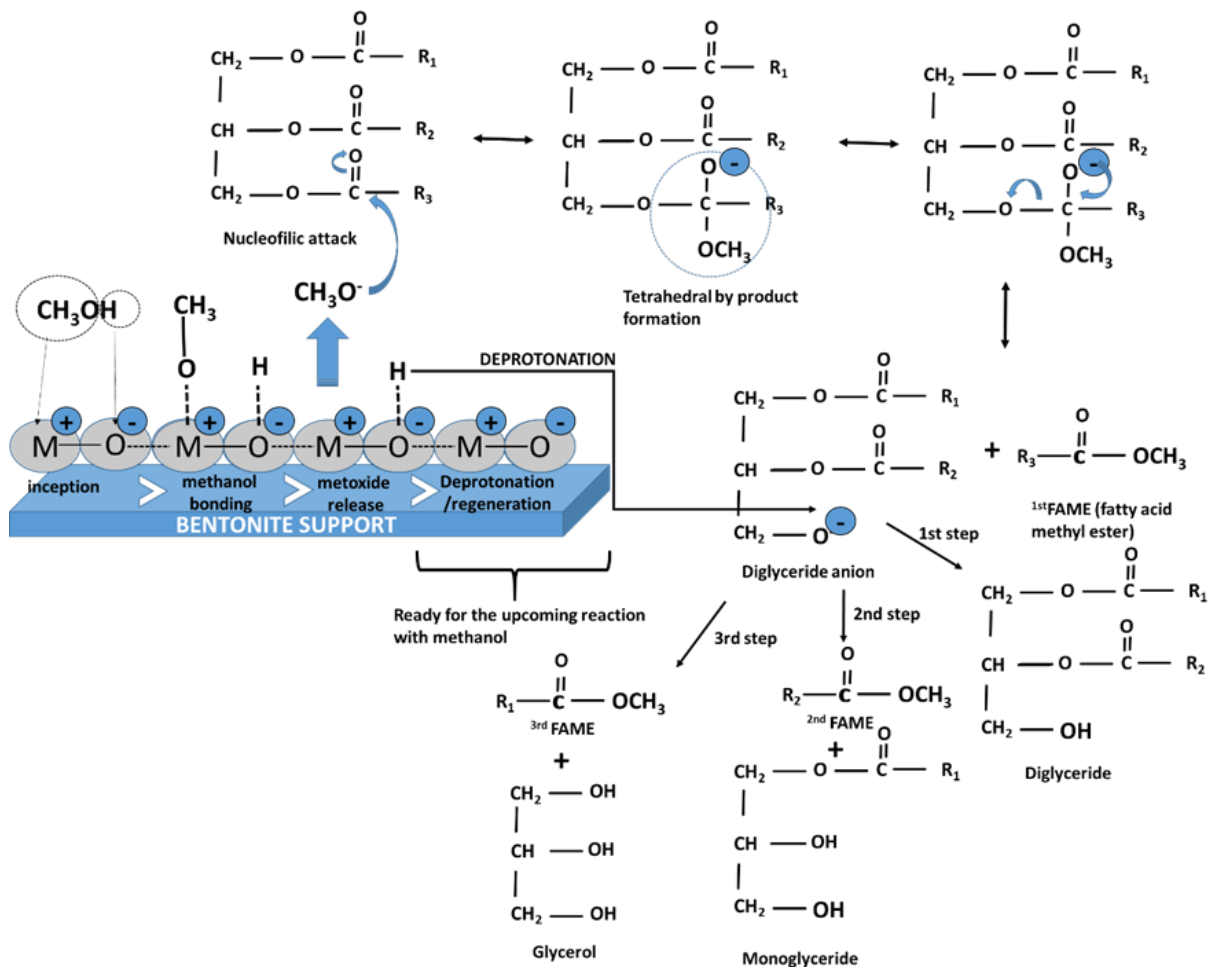
¹ Sema ASLAN (Orcid ID: <https://orcid.org/0000-0001-9796-7311>), Muğla Sıtkı Koçman Üniversitesi, Fen Fakültesi, Kimya Bölümü, Muğla, Türkiye

*Sorumlu Yazar/Corresponding Author: Sema ASLAN, e-mail: semaaslan@mu.edu.tr

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INTRODUCTION

It is clear that pollution levels are reaching higher grades and cause deterioration of the health of the live organisms (Tursi, 2019). Especially petroleum sourced pollution increasing by the usage of fossil fuels for industrial or postindustrial studies. These effects lead the researchers to find out more and more renewable energy sources (Chen, 2017). Thus biomass usage is preferred instead of conventional energy sources (Sani et al., 2014). It should be noted that it is a completely renewable source because it releases biogenic originated CO₂ to the atmosphere and plants use this CO₂ for their growth and metabolic processes (Tkemaladze and Makhashvili, 2016, Sun et al., 2010). As biodiesel is a renewable and environmentally friendly fuel, it shows great potential as an alternative energy source (Guyon and Elisseff, 2003; Rana et al., 2018). It is produced by many divergent methods (microwave-assisted (Moser, 2009; Kumar et al., 2011), sonication (Al-Ani et al., 2018), direct transesterification (Aslan et al., 2019), microbial (Akia et al., 2014) or sources, different kinds of raw materials can be used (refine, crude or frying oils) (Li et al., 2014), or there are different types of catalyst utilization, beyond acids or bases (Karahan et al., 2016; Donmez et al., 2018; Gungor et al., 2018; Kara et al., 2018), photocatalysts (Uğurlu et al., 2006; Karaoğlu et al., 2010; Karaoğlu and Uğurlu, 2010; Dinçer et al., 2016) ion exchange resins, lipases or supercritical fluids are used (Marchetti and Erazzu, 2008; Duan and Savage., 2011). Generally, biodiesel is produced by the transesterification reaction of triglycerides with short chain alcohols in the presence of a suitable catalyst (Izadpanahi et al., 2012).



Scheme 1. Methoxide formation on the BT support and the reaction mechanism of biodiesel formation

Catalyst transesterification reactions are mostly carried out by using alkaline, acidic or enzymatic catalysts (Duan and Savage, 2011). Recently, microporosis heterogeneous catalysts such as zirconium (Erkarlan et al., 2018), sulfated zirconium and titanium based zeolite have been used in the biodiesel production from different oil sources (Haupt and Haupt, 2004). The transesterification reaction takes place at a temperature of 250 °C in the catalyst-free environment but the same reaction can be successfully operated at lower temperatures by the usage of acidic (Suryaputra et al., 2013), alkaline or heterogeneous catalysts (Kennedy and Eberhart, 2001; Poore et al., 2012). Biodiesel production was achieved by a transesterification reaction in the reflux set up. The transesterification reaction is the conversion of triglycerides into fatty acid methyl esters (FAME) and glycerol (Scheme 1).

In recent years, efforts to develop heterogeneous catalysts (Karaoğlu et al., 2012; Dinçer et al., 2016) have been accelerated to use in biodiesel production. Post-process separation, lack of purification processes, eliminating the neutralization step of free fatty acids, no saponification of triglycerides, no catalytic residue in methyl ester and glyceride, and making production more economical are the main reasons for this development (Karaoglu and Ugurlu, 2011; Wang et al., 2019). At the same time, high product conversion can be achieved in heterogeneous catalysts and these eco-friendly catalysts are renewable catalysts (Li and Chung, 2008, Soetaredjo et al., 2011, Sen et al., 2016).

In this study, bentonite (BT) was activated by calcination (C-BT) and Li impregnated bentonite (Li-BT) was produced. Two different basic heterogeneous catalysts were utilized for the biodiesel production from rapeseed oil by transesterification method. For this purpose, the morphological and analytical characteristics of the catalysts and biodiesels were investigated with SEM/EDS, FTIR, TGA and GC-MS analysis. Li-BT revealed the best biodiesel conversion performance due to the presence of Li ions rather than being able to replace H⁺ atoms in order to obtain methoxide from methanol in the first step of biodiesel production process.

MATERIALS AND METHODS

Materials

Methanol (99.80 % pure), n-hexane (99.00 % pure), natrium sulfate (99.00 % pure) and lithium acetate (99.00 % pure) were purchased from Merck. BT was supplied from a local plant. Rapeseed oil was obtained from a local market (Aypa Co.).

Experimental Procedure

Preparation of the heterogeneous catalysts

Firstly natural BT particles were minimized in a mortar and sieved into 350.00 µm. Then obtained particles were activated by calcination at 500.00 °C for 2.00 hours. This form of BT was used as C-BT and a suitable amount of this sample kept for the experiments.

Subsequently, in order to produce Li-BT, 5.88 g of CH₃COOLi_(s) was dissolved in 500.00 mL of deionized pure water in a suitable flask and 20.00 g of C-BT was added to this solution. The obtained mixture was maintained at 60.00 °C and stirred mechanically for 24.00 h. The resulting slurry was transferred to drying oven at 110.00 °C for 24.00 h. Finally calcination was operated at 500.00 °C for 5.00 h. Then, these samples were transferred into a desiccator, cooled and stored in dark medium when not in use (Omar and Amin, 2011).

The morphological and structural properties of the C-BT and Li-BT catalysts were investigated by TGA, FT-IR (Perkin Elmer), SEM (JEOL, JSM-7600F scanning electron microscope) and EDS analysis.

Transesterification and biodiesel production process

Biodiesel is the total FAME content of the mixture. Firstly, methanol (methanol/oil; 6/1) and a small quantity of heterogeneous catalyst (6.00 %) were transferred to the reflux condenser adapted three-neck flask. Heated and stirred at 60.00 °C for 1 h in order to obtain methoxide groups on the heterogeneous catalyst. Subsequently, rapeseed oil was added to the mixture and stirred for 3.00 h at the same conditions. Then the mixture was cooled down to room temperature. Obtained samples were centrifuged for 15.00 min at 5000.00 rpm and the supernatant solution was washed and dried with sodium sulfate anhydrous. Finally, excess methanol was evaporated at 105.00 °C. Total FAME content of the analyzed by GC-MS (Omar and Amin, 2011).

RESULTS AND DISCUSSION

Chemical Characterizations of Catalysts

FTIR spectroscopy analysis

The FTIR spectra of natural BT, C-BT, and Li-BT were analyzed in order to explain structural changes after the calcination and Li impregnation process (Figure 1). There has been a decrease at the peak of natural BT in the field of OH vibration after the calcination (3625.63 cm^{-1}), so as the OH bending of zeolitic water (1636.55 cm^{-1}) depending on the removal of humidity in the natural BT structure. It can be concluded that after calcination of BT, CO bending peak is suppressed but after Li impregnation, the interaction between CO group and Li ions increased the peak (2157.71 cm^{-1} , 2159.99 cm^{-1} , 2160.57 cm^{-1} , respectively). About the fingerprint area, the 1009.16 cm^{-1} peak of the natural BT corresponds to the C-C single bond or Si-O-Si vibrations and as it seems from the spectra this bond is suppressed after the calcination but increased again after the Li impregnation. Finally, the peak at the 793.81 cm^{-1} in the fingerprint area shows the same behavior (Cobas et al., 2014; Eren et al., 2009). These results mean that calcination was achieved successfully and Li interactions were observed especially with C-C and C=O bonds.

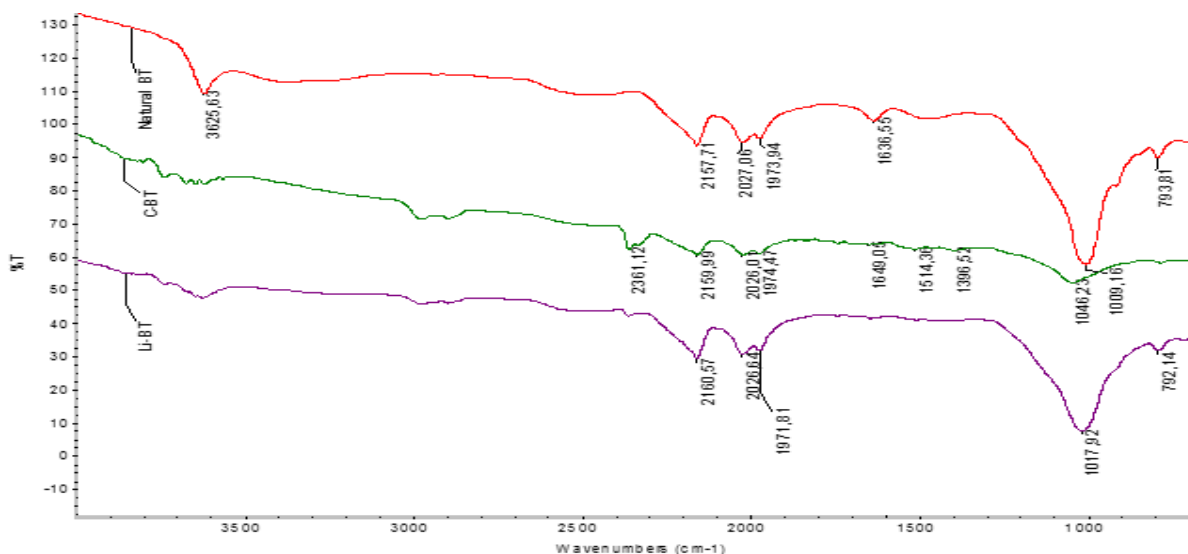


Figure.1. Comparison of the FTIR analysis of natural BT, C-BT, and Li-BT ($4000\text{-}400\text{ cm}^{-1}$)

TGA analysis of catalyst

Thermogravimetric analysis of natural, calcinated and Li impregnated BTs were performed to validate FTIR results. Weight loss percentages of natural BT is higher than C-BT (Figure 2a) and Li-BT (Figure 2b). As it has been mentioned above, there is a water presence in the natural BT structure. So, the weight loss at 100.00 °C is attributed to the water content of the BT and after the calcination, no weight loss is observed in this area for both C-BT and Li-BT samples. Additionally, between 600.00-700.00 °C maximum carbon dioxide release belongs to the natural BT, then C-BT and finally Li-BT. These results can be explained by the presence of Li interaction between C atoms.

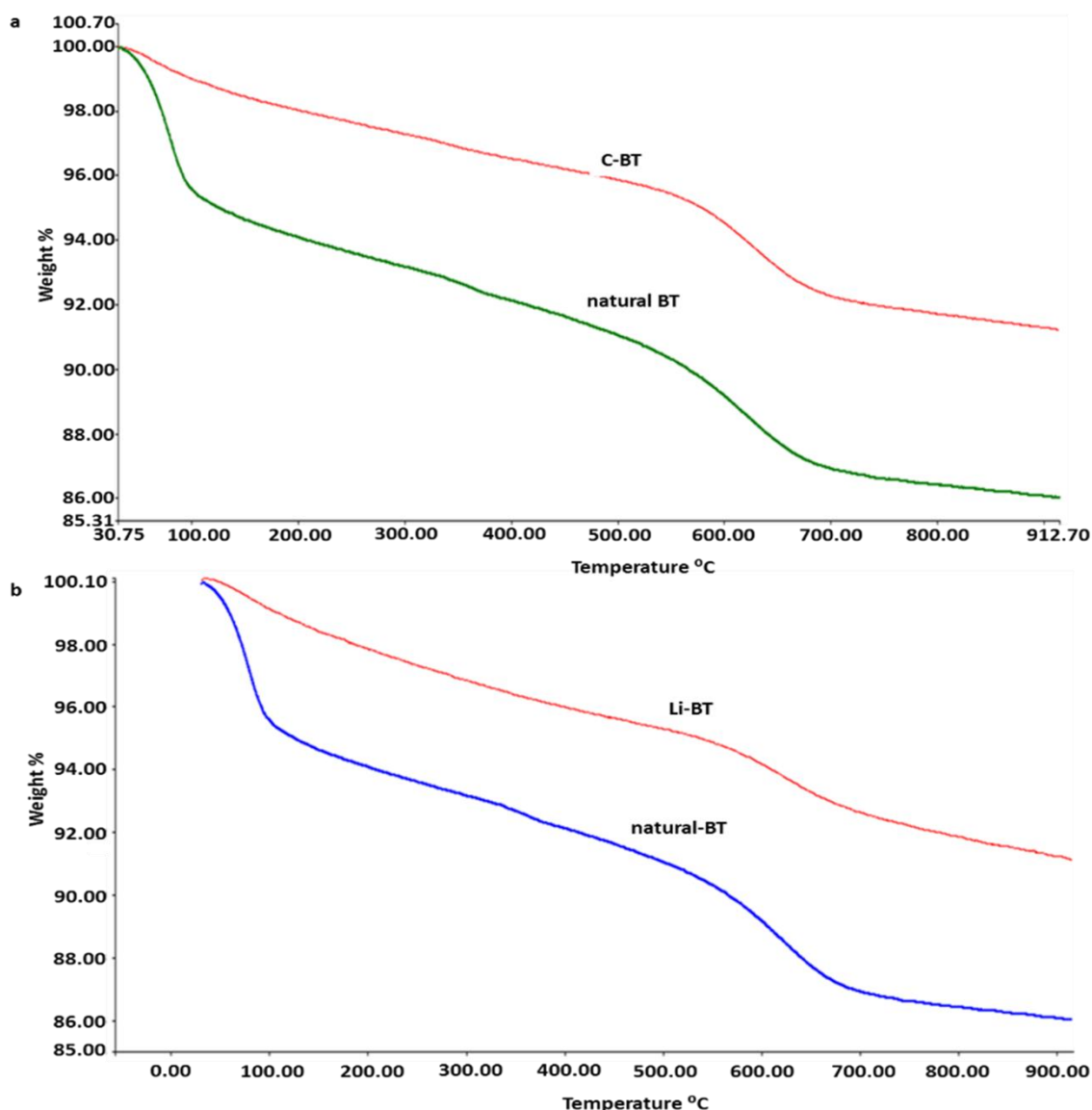


Figure 2. Comparison of the TGA analysis of a) natural BT and C-BT, b) natural BT and Li-BT (21-900 °C)

SEM and EDS analysis of catalysts

The morphological properties of the catalysts were monitored with SEM and EDS analysis. Here the SEM images and EDS values of C-BT and Li-BT are presented and compared with each other. Since Li element shows the EDS peak around 0.02 keV it was expected to be overlapped by C and K peaks in the spectrum additionally if the irradiation beam of the electrode isn't high enough to indicate the Li ion it can't be seen or indicated as another ion here such as K. As it is reported (Hovington et al., 2017)

lithium is highly mobile under the electron beam and therefore any quantitative measurements should be treated with caution. According to the EDS values, the change in the elemental composition of C-BT is attributed to the successful impregnation with Li on the surface of BT. Besides SEM image of the Li-BT shows morphological differences compared to the C-BT surface it can be concluded that Li impregnation was achieved successfully.

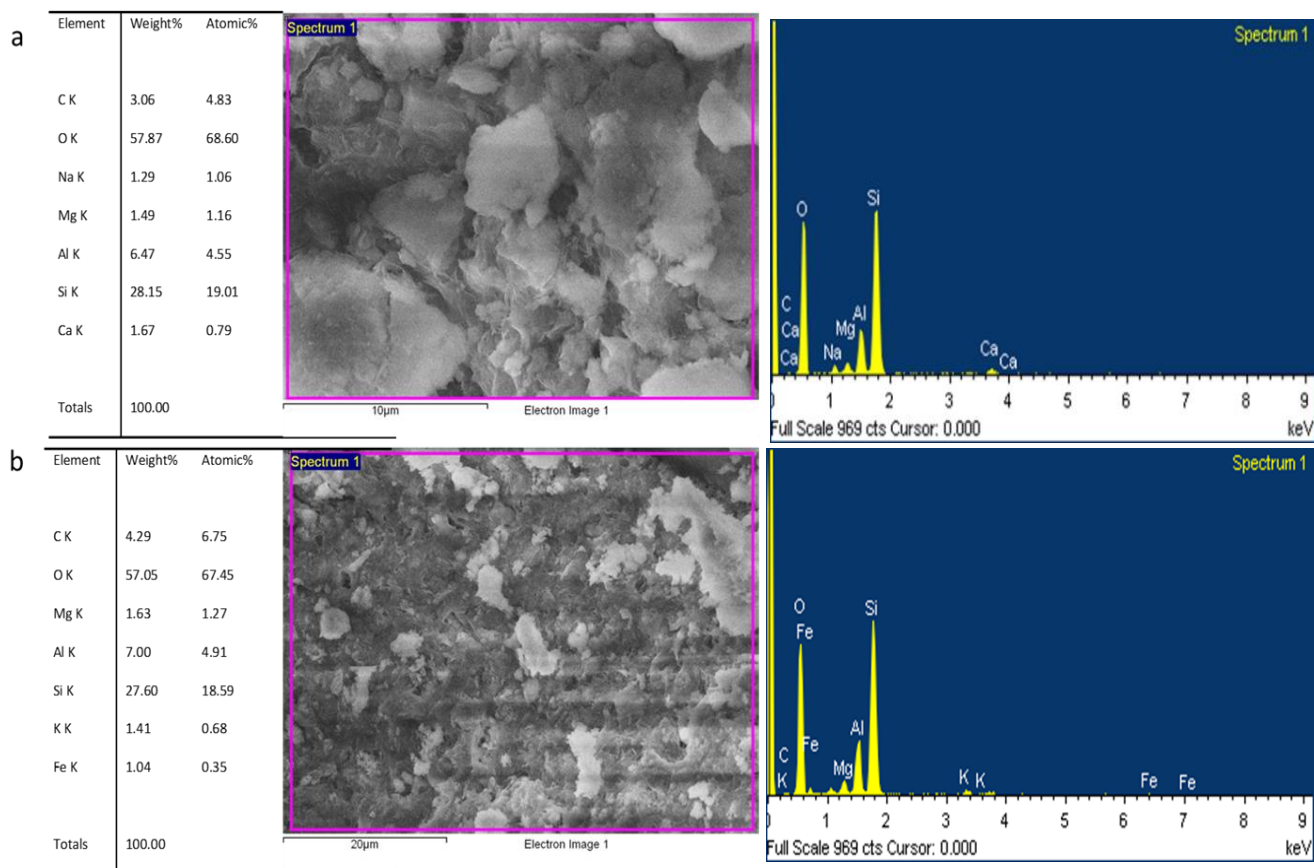


Figure 3. Comparison of the EDS analysis results and SEM imaginations of a) C-BT, b) Li-BT.

Evaluation of the biodiesel production performance of Li-BT.

The fuel properties of biodiesel are indicated according to the FAME composition, along with the presence of other contaminants as water or alcohol and minor components such as sulphur, phosphines or iodide. This hint leads the researchers to explore suitable feedstock because each feedstock has its unique chemical composition. Another crucial point is conversion achievement after the process completed. At this point, GC-MS analysis of produced biodiesel is necessary (Dias et al., 2008). Total FAME content of biodiesel fuel indicates the quality of the components of whole process; beginning from feedstock choice goes on with true catalyst choice and ends with convenient purification. Here the results of the GC-MS analysis (Figure 4) are given in detail in Table 1 with peak area percentages of corresponding FAME types.

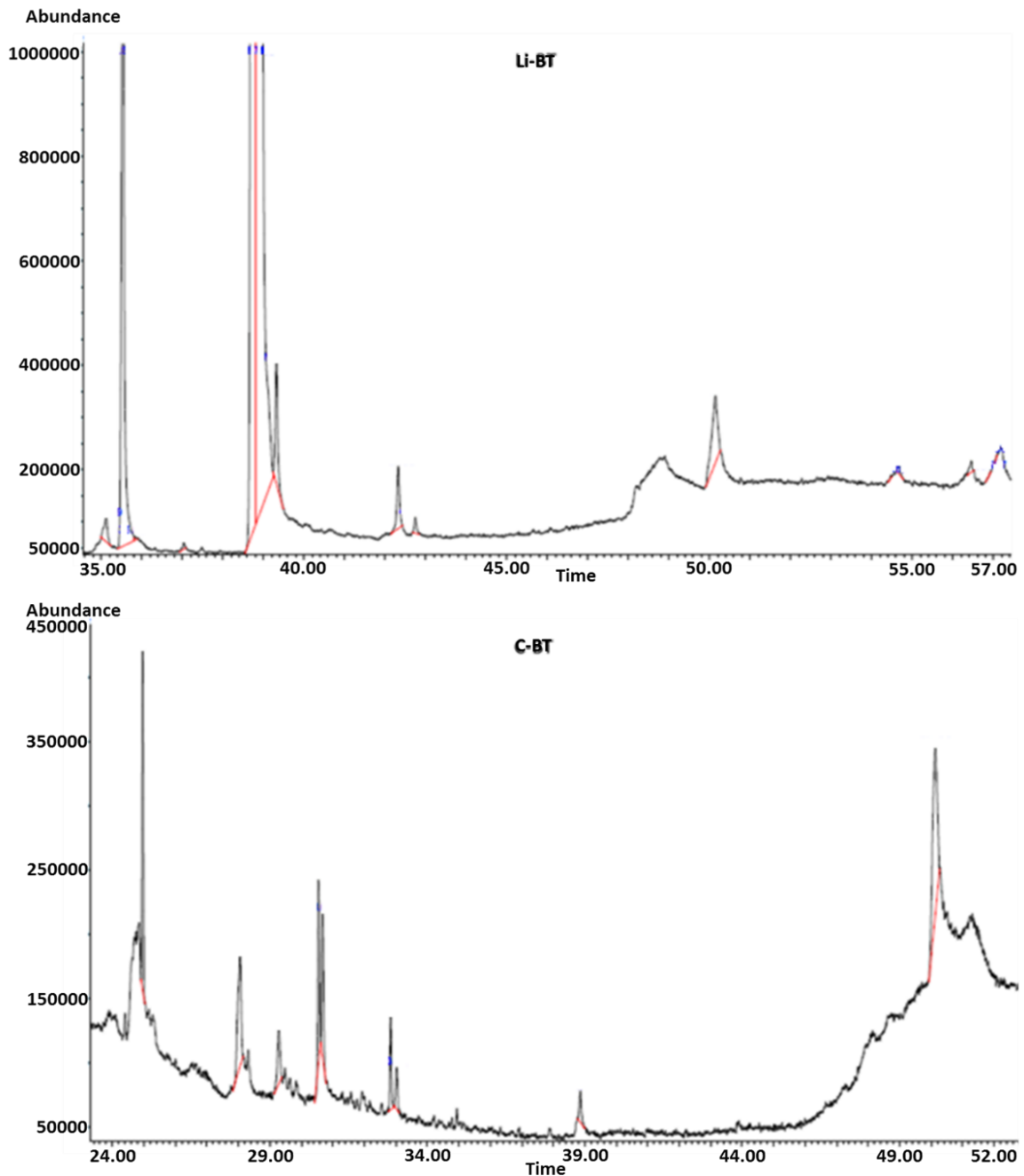


Figure 4. GC-MS spectra of the biodiesel products of Li-BT and C-BT catalysts.

Table 1. FAME composition of the final biodiesel product of Li-BT catalysed process

FAME type	Unit quantity of the product (peak area %)
Methyl palmitate hexadecanoic acid	3.86
6-Methoxybenzofuroxan	0.16
Methyl linoleate	30.84
Oleic acid methyl ester	63.31
Methyl stearate	0.56
cis-11-Eicosenoic acid methyl ester	0.06
Arsenous acid, tris(trimethylsilylester)	0.01

Another important biodiesel quality determination factor is the carbon to hydrogen ratio. According to the heat of combustion, values FAME have similar chain length can show different energy outputs. If they have lower carbon to hydrogen ratios, mainly containing more hydrogen than carbon, they exhibit greater energy values. For instance, stearic acid methyl esters (largest hydrogen content) show greater energy yields than oleic, linoleic, and linolenic (smallest hydrogen content) acids [12]. Therefore, lower energy potential is obtained from regularly higher levels of unsaturation with otherwise similar chain length (DeOliveira et al., 2006). Obtained results indicate the expected energy value of the presented biodiesel is low but the conversion percentage is very satisfactory. Because the oleic and linoleic acid methyl ester amounts are major elements in the final fuel (30.84 and 63.31 % respectively) C to H ratio of the fuel is expected to be low.

CONCLUSION

The choice of the feedstock is a very crucial point to determine final fuel FAME quality and so as the alcohol type which is used for esterification reaction. Besides, homogeneous catalyst usage results in saponification that is lethal for the transesterification process. The most important point to be emphasized here is the use of BT as a highly economical and effective heterogeneous catalyst. Biodiesel conversion is very satisfactory but the quality of final fuel ingredients can be lifted by the choice of stearic acid rich feedstock choice. Also, the reflux method is very conventional but this process is ready to try with microwave assisted system with shorter reaction times.

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